

	Half Reaction		Standard Potential (V)
	$\text{F}_2 + 2\text{e}^- \rightleftharpoons 2\text{F}^-$		+2.87
	$\text{Pb}^{4+} + 2\text{e}^- \rightleftharpoons \text{Pb}^{2+}$		+1.67
	$\text{Cl}_2 + 2\text{e}^- \rightleftharpoons 2\text{Cl}^-$		+1.36
	$\text{O}_2 + 4\text{H}^+ + 4\text{e}^- \rightleftharpoons 2\text{H}_2\text{O}$		+1.23
	$\text{Ag}^+ + \text{e}^- \rightleftharpoons \text{Ag}$		+0.80
	$\text{Fe}^{3+} + \text{e}^- \rightleftharpoons \text{Fe}^{2+}$		+0.77
	$\text{Cu}^{2+} + 2\text{e}^- \rightleftharpoons \text{Cu}$		+0.34
	$2\text{H}^+ + 2\text{e}^- \rightleftharpoons \text{H}_2$		0.00
	$\text{Pb}^{2+} + 2\text{e}^- \rightleftharpoons \text{Pb}$		-0.13
	$\text{Fe}^{2+} + 2\text{e}^- \rightleftharpoons \text{Fe}$		-0.44
	$\text{Zn}^{2+} + 2\text{e}^- \rightleftharpoons \text{Zn}$		-0.76
	$\text{Al}^{3+} + 3\text{e}^- \rightleftharpoons \text{Al}$		-1.66
	$\text{Mg}^{2+} + 2\text{e}^- \rightleftharpoons \text{Mg}$		-2.36
	$\text{Li}^+ + \text{e}^- \rightleftharpoons \text{Li}$		-3.05

stronger oxidizing agent ↑

↓ stronger reducing agent

# Electrode Potentials

**SA Adler**

## **Electrode Potentials:**

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dedicado a su música en ese momento el joven comenzó a componer y sus primeras obras fueron interpretadas por una orquesta de alumnos de a s internado de la ciudad donde fue violinista en 1813 se jubila para empezar a trabajar como

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