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ELEMENTS

of Physical Chemistry

PETER ATKINS & JULIO DE PAULA



7th edition

Elements Of Physical Chemistry

John Livingston Rutgers Morgan



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$$\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$$

$$2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$$

$$3\text{HBr} + 1\text{KHCO}_3 + 1\text{H}_2\text{O} + 1\text{KBr} + 1\text{CO}_2$$

$$2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 3\text{H}_2\text{SO}_4 + 2\text{H}_2\text{O} + 2\text{Na}_2\text{SO}_4$$
br using the following equation

$$\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$$

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$$\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$$

$$2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$$

$$\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$$

$$2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 3\text{H}_2\text{SO}_4 + 2\text{H}_2\text{O} + 2\text{Na}_2\text{SO}_4$$

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$$\text{SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$$

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$$\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$$

$$2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$$

$$\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$$

$$2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 3\text{H}_2\text{SO}_4 + 2\text{H}_2\text{O} + 2\text{Na}_2\text{SO}_4$$

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