

THE ELECTRONIC THEORY OF VALENCY.

A GENERAL DISCUSSION.

A GENERAL DISCUSSION on "THE ELECTRONIC THEORY OF VALENCY" was held by The Faraday Society on July 13th and 14th, 1923, in the Department of Physical Chemistry, University of Cambridge.

PART I.—GENERAL.

Friday, July 13th, 1923.

Chairman:—Sir J. J. Thomson, O.M., F.R.S., Master of Trinity College.

The **Chairman** opened the proceedings with the following remarks:

The electron dominates the field of chemistry, which finds its most suggestive mode of expression in terms of electrons. Admitting the presence of electrons, their repulsion involves important chemical consequences. The tendency of electrons is to scatter, and they are only kept together by the positive charge on the nucleus. The essence of the question is that, for a number of electrons arranged at equal distances, the disruptive force increases rapidly—as third power of the number present—while the attraction of the nucleus is only proportional to its charge. Thus a limit is set to the number of electrons. The exact number depends upon the nature of the forces acting upon them. In practice electrons make layers of 8; but the law of force might have been different, and we might have had sextets instead. These phenomena were first experimentally investigated by Meyer in America with floating magnets, but the method had been extended by the speaker.

It was desirable to pay attention to the forces brought about by different distributions of electrons, *e.g.* what alteration was produced by change in a substituent? Replace H in CH_4 by Cl. What change in electric force is produced? This is a question which would repay investigation. The speaker on calculating to a first approximation obtained the simple result that an electric moment is introduced into the molecule such as would be produced by a positive and negative charge at a distance given by the sum of the radii of the atoms. Whenever H is replaced by Cl the same electrical doublet should be introduced. These moments can be calculated per molecule from measurements of specific inductive capacity

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