

Thermodynamics

James H. Luscombe



A diagram illustrating the relationship between the heat differential dQ/T and the entropy differential $dS_i \geq 0$. The expression dQ/T is written in orange above a double-headed vertical arrow. The arrow points down to a rectangular box with an orange border. Inside the box, the expression $dS_i \geq 0$ is written in orange.

$$dQ/T$$
$$dS_i \geq 0$$



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Elementary Thermodynamics 1st Edition

Lauren Gardner



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Basic Chemical Thermodynamics (Fifth Edition) E Brian Smith, 2004-04-08 This widely acclaimed text now in its fifth edition and translated into many languages continues to present a clear simple and concise introduction to chemical thermodynamics An examination of equilibrium in the everyday world of mechanical objects provides the starting point for an accessible account of the factors that determine equilibrium in chemical systems This straightforward approach leads students to a thorough understanding of the basic principles of thermodynamics which are then applied to a wide range of physico chemical systems The book also discusses the problems of non ideal solutions and the concept of activity and provides an introduction to the molecular basis of thermodynamics Over five editions the views of teachers of the subject and their students have been incorporated The result is a little more rigour in specifying the dimensions within logarithmic expressions the addition of more worked examples and the inclusion of a simple treatment of the molecular basis of thermodynamics Students on courses in thermodynamics will continue to find this popular book an excellent introductory text a

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Solution Thermodynamics and Its Application to Aqueous Solutions Yoshikata Koga, 2017-03-28 Solution Thermodynamics and its Application to Aqueous Solutions A Differential Approach Second Edition introduces a differential approach to solution thermodynamics applying it to the study of aqueous solutions This valuable approach reveals the molecular processes in solutions in greater depth than that gained by spectroscopic and other methods The book clarifies what a hydrophobe or a hydrophile and in turn an amphiphile does to H₂O By applying the same methodology to ions that have been ranked by the Hofmeister series the author shows that the kosmotropes are either hydrophobes or hydration centers and that chaotropes are hydrophiles This unique approach and important updates make the new edition a must have reference for those active in solution chemistry Unique differential approach to solution thermodynamics allows for experimental evaluation of the intermolecular interaction Incorporates research findings from over 40 articles published since the previous edition Numerical or graphical evaluation and direct experimental determination of third derivatives enthalpic and volumetric AL AL interactions and amphiphiles are new to this edition Features new chapters on spectroscopic study in aqueous solutions as well as environmentally friendly and hostile water aqueous solutions

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