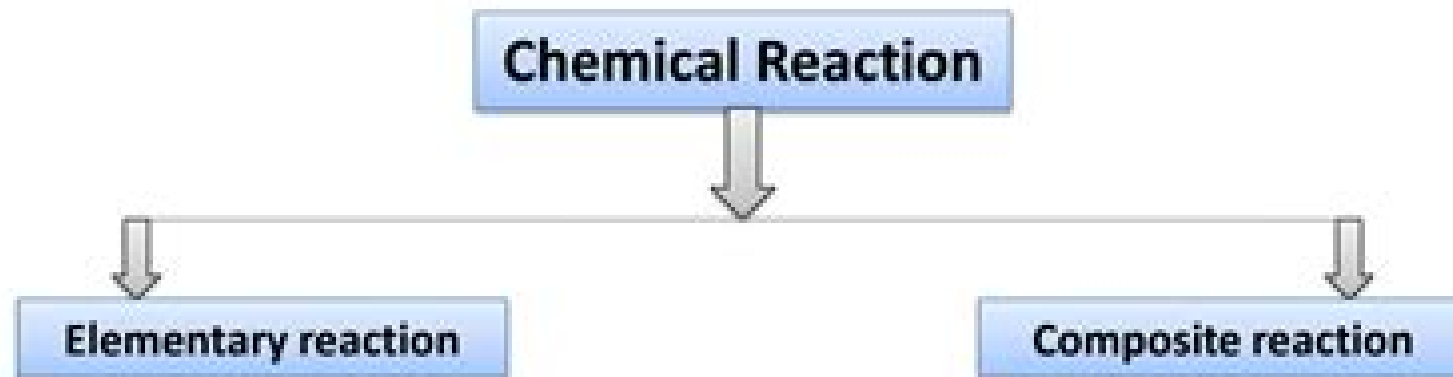




An **elementary reaction** is a simple reaction which occurs in a single step.

Example: $A \longrightarrow \text{product}$



A **composite or complex reaction** is that which occurs in two or more steps.

Molecularity of an elementary reaction

The **molecularity** of an elementary reaction is defined as the number of reactant molecules involved in a reaction.

The elementary reactions having molecularity 1, 2 and 3 are called **unimolecular, bimolecular and termolecular reaction** respectively.

Elementary Mechanism

Robert R. Redfield



Elementary Mechanism:

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