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Sailing Directions for the English Channel Anonymous, 2023-03-04 Reprint of the original first published in 1872

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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calculated mass of the product expected x100 in this reaction a yield 80 is anticipated

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g Al x 1 mol Al x 2 mol Al_2O_3 x 101 964 g Al_2O_3 302 3 g Al_2O_3 26 982 g Al 4 mol Al 1 mol Al_2O_3 c determine the percent yield
yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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web 7 Na 2 C 2 O 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be
weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12
C is the limiting reagent 4 33 g of H 2 are left over 13

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent
yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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CH₃OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH₃OH if 68 5 g of CO is
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reaction

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 $80 \text{ g I}_2\text{O}_5 \times \frac{1 \text{ mol I}_2\text{O}_5}{1 \text{ mol I}_2} \times \frac{1 \text{ mol I}_2}{1 \text{ mol I}_2\text{O}_5} \times \frac{253.8 \text{ g I}_2}{1 \text{ mol I}_2} = 333.8 \text{ g I}_2$
 $28 \text{ g CO} \times \frac{1 \text{ mol CO}}{28 \text{ g CO}} \times \frac{1 \text{ mol I}_2}{1 \text{ mol CO}} = 1 \text{ mol I}_2$

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage
 $\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100$ label 3 7 3