

# ATOMIC STRUCTURE

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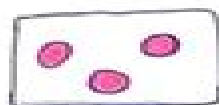
**ATOM** - The basic building block for all matter in the universe. Single particle.

**PROTON** - A positively charged particle that is found in the nucleus of the atom.

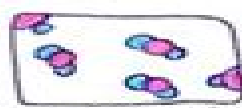
**NEUTRON** - A particle found in the nucleus with no charge.

**ELECTRON** - A negatively charged particle that spins around the outside of the nucleus.

**ELEMENT** - A substance consisting of atoms which all have the same number of protons.



ELEMENT



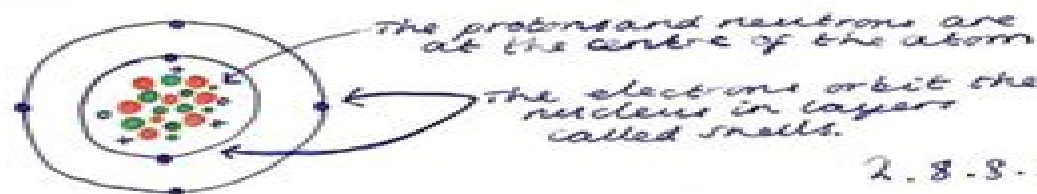
COMPOUND



MIXTURE

**COMPOUND** - 2 or more different elements put together, chemically joined together.

**MIXTURE** - combination of two or more pure substances.



- **ATOMIC NUMBER** = number of protons.

Atomic mass - neutron = proton / electron.

2.8.8.13

## THE LAW OF CONSERVATION OF MASS

The total mass of the products formed in a reaction is equal to the total mass of the reaction.

[Mass number = number of protons + number of neutrons.]

**IONS** - an ion is an atom or group of atoms that have an electrical charge, either positive or negative.

• Elements can gain or lose electrons during a chemical reaction.

• **POSITIVE ION** - an atom that loses electrons has more protons than electrons and so has a positive overall charge.

• **NEGATIVE ION** - an atom that gains electrons has more electrons than protons and so has a negative overall charge.

- protons don't change.

• **METALS** tend to [lose] electrons.

• **NON-METALS** tend to [gain] electrons.

• **ISOTOPES** - are atoms that have the same number of protons but different numbers of neutrons in their nuclei.

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