

KINETICS AND TRANSPORT IN ELECTRODE REACTIONS

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6.1 Introduction

In the last two chapters the kinetics of electrode processes and mass transport to an electrode were discussed. In this chapter these two parts of the electrode process are combined and we see how the relative rates of the kinetics and transport cause the behaviour of electrochemical systems to vary¹⁻⁴.

6.2 The global electrode process: kinetics and transport

Transport to the electrode surface as described in Chapter 5 assumes that this occurs solely and always by diffusion. In hydrodynamic systems, forced convection increases the flux of species that reach a point corresponding to the thickness of the diffusion layer from the electrode. The *mass transfer coefficient* k_d describes the rate of diffusion within the diffusion layer and k_c and k_a are the *rate constants* of the electrode reaction for reduction and oxidation respectively. Thus for the simple electrode reaction $O + ne^- \rightarrow R$, without complications from adsorption,

Encyclopedia Of Electrochemistry Interfacial Kinetics And Mass Transport

M Tight



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