

De enzymatische analyse

Rol van enzymen

- Enzymen dienen als biologische katalysators
 - Ze versnellen biochemische reacties, zodat deze niet oneindig traag zouden verlopen
 - Elke metabole pathway kent zijn eigen biologische katalysator
 - hoge specificiteit voor het binden van substraat
 - Verschillend aan chemische katalysatoren
 - Meestal opgebouwd uit eiwitten, soms in combinatie met cofactoren
 - Gaan niet deelnemen aan de biochemische reactie
 - na de reactie is het enzyme onveranderd gebleven (niet verbruikt)
- Enzymen kunnen onder mildere condities werken bij biochemisch dan bij chemische reacties
 - Geen hoge druk of hoge temperaturen nodig

Begrippen van de thermodynamica

- Cellen hebben een zeer efficiënt mechanisme ontworpen om energie uit nutriënten te halen via katabole pathways en nadien te converteren naar een vorm van energie dat gebruikt kan worden in de anabole pathways
- Een chemische reactie kan spontaan en niet-spontaan verlopen
 - Gibbs-vrije energie geeft hier een indicatie van de spontaniteit
 - $\Delta G = \Delta H - T \Delta S$
 - Met H = enthalpie en S = entropie
 - Indien $\Delta G > 0$ → product heeft een hogere energie dan de reagentia
 - Endergonische reactie (niet-spontaan)
 - Er is energie nodig om de reactie door te laten gaan
 - Indien $\Delta G < 0$ → product heeft lagere energie dan de reagentia
 - Exergonische reactie (spontaan)
- Op een gegeven moment bereikt de reactie een evenwicht → $\Delta G = 0$
 - Constante kan gedefinieerd worden = K_{eq} = evenwichtsconstante
 - $\frac{[C]_{eq} \cdot [D]_{eq}}{[A]_{eq} \cdot [B]_{eq}} = K_{eq}$
 - Hoe groter K, hoe meer het evenwicht naar de producten zal liggen
 - Meer reagentia omgezet naar producten tijdens chemisch evenwicht
 - Constante voor elke chemische reactie
- De Gibbs-vrije energie is afhankelijk van de initiële concentraties van reagentia en producten
 - Hoe meer reagentia aanwezig is, hoe meer product er kan gevormd worden
 - Hoe meer product er gevormd is, hoe lager ΔG tot bereiken van evenwicht
 - $\Delta G = \Delta G^0 + R \cdot T \cdot \ln \frac{[C] \cdot [D]}{[A] \cdot [B]}$
 - Met ΔG^0 = standaard vrije energie verandering (pH = 7, atm en 25°C)
 - Constante voor elke chemische reactie
- Hoe groter het evenwicht, hoe lager de ΔG^0 → spontane reactie

Enzymatische Analyse

G Orfield



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