EXAMPLE 16-9

$$H_2PO_4^- + H_2O \rightleftharpoons HPO_4^{2-} + H_3O^+$$

Initial conc.

0.14 M

0

0.14 M

Changes

-y M

+y M

+y M

Equilibrium

(0.14 - y) M

y M

(0.14 + y) M

Concentration

$$K_{a} = \frac{[H_{3}O^{+}][HPO_{4}^{2-}]}{[H_{2}PO_{4}^{-}]} = \frac{y \cdot (0.14 + y)}{(0.14 - y)} = 6.3 \times 10^{-8}$$

$$y \le 0.14 \text{ M}$$

$$y = [HPO_4^{2-}] = 6.3 \times 10^{-8}$$

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