

General Chemistry 101 Lecture

Chapter 4: Reactions in Aqueous Solution

4.1: General Properties of Aqueous Solutions

- A solution is a homogeneous mixture of 2 or more substances
 - The **SOLVENT** is the substance w/ the greatest quantity
 - When water is the solvent, the solution is called an aqueous solution
 - The **SOLUTES** are the other substance(s)
 - They're dissolved in the solvent
- An **ELECTROLYTE** is a substance whose aqueous solutions contain ions
 - Dissociates into ions when dissolved in water
 - A **STRONG ELECTROLYTE** dissociates completely when dissolved in water
 - A **WEAK ELECTROLYTE** only partially dissociates when dissolved in water
- A **NONELECTROLYTE** is a substance that doesn't form ions in solution
 - May dissolve in water, but it doesn't dissociate into ions when dissolved
 - DOESNT dissociate in water
- **SOLVATION** helps stabilize the ions in solution & prevents cations & anions from recombining
- **DISSOCIATION** occurs when an ionic compound dissolves in, the ions become surrounded in H₂O molecules
 - Helps stabilize the ions in solution & prevents cations & anions from recombining
 - Most molecular compounds DONT dissociate
- Reactions who are significant in both directions have a balance that produces a state of **CHEMICAL EQUILIBRIUM**
 - Chemists use half-arrows pointing in opposite directions to represent reactions that go forward & backward to achieve equilibrium
 - Example: ionization of weak electrolytes

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Anthony S. Fauci



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