Gen Chem 2 LG Session 2

Answer Key

Question 1: A 1.12 M solution of citric acid (H3C6H5O7) in water has a density of 1.08 g/cm3. Calculate the mass percent, molality, and mole fraction of the citric acid.

Assume 1 L of Solution: 12 moles (the Acid & mass of Molar Mass is 192,123 & Solution is 1080 grams

Molar Mass is 192,123 & Mass % = 215.178 grams

Mass of 1420 = 1089 - 215.178 grams

Molality = 1.12 moles = 1.3 mol mg

Molality = 1.12 moles = 1.3 moles = 1.12 moles = 1.12 mg

Molality = 0.86482 mg

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Molality = 1.12 moles = 1.3 moles = 1.

Question 2: Pure solvent X boils at 81.1°C. When 11.00 g of solute A (non-electrolyte, MM = 88.1 g/mol) is dissolved in 150.0 g of solvent X, the new boiling point is 83.5°C. When 4.00 g of solute B (non-electrolyte) is dissolved in 50.0 g of solvent X, the solution boils at 82.4°C. What is the molar mass of solute B?

A:
$$\Delta B_{\rho}^{-} = 2.4^{\circ}C$$

VanHff = 1

Molality = $\frac{\text{Moles}}{\text{Rig Solvent}} = \frac{(13/38.19/mol)}{0.150 \text{ kg}} = 0.8324 \text{ mol}$
 $2.4 = \Delta B_{\rho} = 1.0.8324 \text{ mg} \cdot \text{Kb}$
 $1 \times b = 2.4^{\circ}C$
 $1 \times 8324 \text{ kg mol}$
 $1 \times b = 2.4^{\circ}C$
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