

## PROBLEMS

- 6.1 Calculate  $\Delta H_{1000}$  and  $\Delta S_{1000}$  for the reaction  $\text{Zr}_{(\text{gr})} + \text{O}_2 = \text{ZrO}_{2(\text{gr})}$ .  
 6.2 Which of the following two reactions is the more exothermic?

1.  $\text{C}_{(\text{graphite})} + \frac{1}{2} \text{O}_{2(\text{g})} \rightarrow \text{CO}_{(\text{g})}$  at 1000 K
2.  $\text{C}_{(\text{diamond})} + \frac{1}{2} \text{O}_{2(\text{g})} \rightarrow \text{CO}_{(\text{g})}$  at 1000 K

- 6.3 Calculate the change in enthalpy and the change in entropy at 1000 K for the reaction  $\text{CaO}_{(\text{s})} + \text{TiO}_{2(\text{s})} \rightarrow \text{CaTiO}_{3(\text{s})}$ .  
 6.4 Copper exists in the state  $T = 298 \text{ K}$ ,  $P = 1 \text{ atm}$ . Calculate the temperature to which the copper must be raised at 1 atm pressure to cause the same increase in molar enthalpy as is caused by increasing its pressure to 1000 atm at 298 K. The molar volume of Cu at 298 K is  $7.09 \text{ cm}^3$ , and the thermal expansivity is  $0.493 \times 10^{-3} \text{ K}^{-1}$ . These values can be taken as being independent of pressure in the range 1–1000 atm.  
 6.5 Calculate  $\Delta H_{298}$  and  $\Delta S_{298}$  for the following reactions:

- a.  $2 \text{TiO} + \frac{1}{2} \text{O}_2 = \text{Ti}_2\text{O}_3$
- b.  $3 \text{Ti}_2\text{O}_3 + \frac{1}{2} \text{O}_2 = 2 \text{Ti}_3\text{O}_5$
- c.  $\text{Ti}_2\text{O}_3 + \frac{1}{2} \text{O}_2 = 3 \text{TiO}_2$

- 6.6 An adiabatic vessel contains 1000 g of liquid aluminum at  $700^\circ\text{C}$ . Calculate the mass of  $\text{Cr}_2\text{O}_3$  which, when added to the liquid aluminum (with which it reacts to form Cr and  $\text{Al}_2\text{O}_3$ ) raises the temperature of the resulting mixture of  $\text{Al}_2\text{O}_3$ ,  $\text{Cr}_2\text{O}_3$ , and Cr to 1600 K.  
 6.7 Calculate the adiabatic flame temperature attained when methane, at 298 K, is combusted (a) with oxygen in the molar ratio  $\text{O}_2/\text{CH}_4 = 2.0$ , and (b) with air in the molar ratio  $\text{air}/\text{CH}_4 = 9.524$ . Assume that  $\text{CO}_2$  and  $\text{H}_2\text{O}$  are the products of combustion. The adiabatic flame temperature is that temperature reached if all of the heat of the oxidation reaction is used to increase the temperature of the products of the reaction. Air is 21 molar percent  $\text{O}_2$  and 79 molar percent  $\text{N}_2$ .  
 6.8 Calculate the value of  $\Delta G$  for the reaction



at 800 K. What percentage error occurs if it is assumed that  $\Delta c_p$  for the reaction is zero?

- 6.9 Determine the stoichiometric coefficients for the reaction



and calculate  $\Delta H_{298}$ ,  $\Delta S_{298}$ , and  $\Delta G_{298}$  for the reaction.

- 6.10 How much heat is required to increase the temperature of 1 kg of cordierite,  $2\text{MgO} \cdot 2\text{Al}_2\text{O}_3 \cdot 5\text{SiO}_2$ , from 298 K to its incongruent melting temperature of 1738 K?

# General Chemistry W Cd Sol Man Math Tlkit

**David Baud**



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