



# Good Science Roof Bks

**Peter James McGuire, Frank Duffy**



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$$\begin{array}{l} 1 \text{ N}_2 + \text{F}_2 \rightarrow 2 \text{ NF}_3 \\ 2 \text{ C}_6\text{H}_{10} + \text{O}_2 \rightarrow 2 \text{ CO}_2 + 2 \text{ H}_2\text{O} \\ 3 \text{ HBr} + \text{KHCO}_3 + \text{H}_2\text{O} \rightarrow \text{KBr} + \text{CO}_2 + \text{GaBr}_3 + \text{NaBr} + \text{SnO}_4 \\ 4 \text{ NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3 \text{ NF}_3 + 2 \text{ SnF}_2 + \text{N}_2\text{O}_3 \end{array}$$
 using the equation from problem 2 above answer the following

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$$\begin{array}{l} 1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3 \\ 2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O} \\ 3 \text{ HBr} + 1 \text{ KHCO}_3 + 1 \text{ H}_2\text{O} \rightarrow 1 \text{ KBr} + 1 \text{ CO}_2 + 2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 + 1 \text{ Ga}_2\text{SO}_3 + 3 \text{ Br}_2 + 3 \text{ SnO}_2 + 3 \text{ NF}_3 + 2 \text{ SnF}_2 + 1 \text{ N}_2\text{O}_3 \end{array}$$
 using the following equation  $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

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$$\text{Ce} + \text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu} + \text{Zn}$$
 how many moles of  $\text{Ce}$  will be produced from 23.0 g of  $\text{Ce}$  assuming  $\text{CuCl}_2$  is available in excess

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$$\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$$

$$\text{C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

$$\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{GaBr}_3 + \text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + \text{NaBr}$$

$$\text{SNO}_2 + \text{NF}_3 \rightarrow \text{SNF}_2 + \text{N}_2\text{O}_3$$
 using the following equation  

$$\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$$

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$$\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{GaBr}_3 + \text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + \text{NaBr}$$

$$\text{SNO}_2 + \text{NF}_3 \rightarrow \text{SNF}_2 + \text{N}_2\text{O}_3$$

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$$\text{SiO}_2 + \text{C} \rightarrow \text{SiC} + \text{CO}$$

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$$\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{GaBr}_3 + \text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + \text{NaBr}$$
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web google classroom balance the following chemical equation  $\text{Mg} + \text{OH}^- \rightarrow \text{Mg}^{2+} + \text{OH}^-$   $\text{HCl} \rightarrow \text{HCl}$   $\text{MgCl}_2 \rightarrow \text{MgCl}_2$   $\text{H}_2\text{O} \rightarrow \text{H}_2\text{O}$  note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem 7 4 1 x x y y theta  $\theta$  pi  $\pi$  8 5 2 0 9 6

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web a common type of stoichiometric relationship is the mole ratio which relates the amounts in moles of any two substances in a chemical reaction we can write a mole ratio for a pair of substances by looking at the coefficients in front of each species in

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