

# Periodic Chart of the Elements

PERIODIC TABLE OF THE ELEMENTS																		NOBLE GASES				
IA																		He				
1	<div>1</div> <div>H</div> <div>1.008</div>																	2	<div>2</div> <div>He</div> <div>4.003</div>			
IIA																		VIIIB				
2	<div>3</div> <div>Li</div> <div>6.941</div>	<div>4</div> <div>Be</div> <div>9.012</div>															<div>5</div> <div>B</div> <div>10.811</div>	<div>6</div> <div>C</div> <div>12.011</div>	<div>7</div> <div>N</div> <div>14.007</div>	<div>8</div> <div>O</div> <div>15.999</div>	<div>9</div> <div>F</div> <div>18.998</div>	<div>10</div> <div>Ne</div> <div>20.179</div>
3	<div>11</div> <div>Na</div> <div>22.990</div>	<div>12</div> <div>Mg</div> <div>24.305</div>															<div>13</div> <div>Al</div> <div>26.982</div>	<div>14</div> <div>Si</div> <div>28.086</div>	<div>15</div> <div>P</div> <div>30.974</div>	<div>16</div> <div>S</div> <div>32.06</div>	<div>17</div> <div>Cl</div> <div>35.45</div>	<div>18</div> <div>Ar</div> <div>39.948</div>
IIIA		IVA		VA		VIA		VIIA		VIII		IB		IIB								
4	<div>19</div> <div>K</div> <div>39.098</div>	<div>20</div> <div>Ca</div> <div>40.078</div>	<div>21</div> <div>Sc</div> <div>44.956</div>	<div>22</div> <div>Ti</div> <div>47.88</div>	<div>23</div> <div>V</div> <div>50.942</div>	<div>24</div> <div>Cr</div> <div>52.00</div>	<div>25</div> <div>Mn</div> <div>54.938</div>	<div>26</div> <div>Fe</div> <div>55.845</div>	<div>27</div> <div>Co</div> <div>58.933</div>	<div>28</div> <div>Ni</div> <div>58.69</div>	<div>29</div> <div>Cu</div> <div>63.546</div>	<div>30</div> <div>Zn</div> <div>65.38</div>	<div>31</div> <div>Ga</div> <div>69.723</div>	<div>32</div> <div>Ge</div> <div>72.64</div>	<div>33</div> <div>As</div> <div>74.922</div>	<div>34</div> <div>Se</div> <div>78.96</div>	<div>35</div> <div>Br</div> <div>79.904</div>	<div>36</div> <div>Kr</div> <div>83.80</div>				
5	<div>37</div> <div>Rb</div> <div>85.468</div>	<div>38</div> <div>Sr</div> <div>87.62</div>	<div>39</div> <div>Y</div> <div>88.906</div>	<div>40</div> <div>Zr</div> <div>91.224</div>	<div>41</div> <div>Nb</div> <div>92.906</div>	<div>42</div> <div>Mo</div> <div>95.94</div>	<div>43</div> <div>Tc</div> <div>98.906</div>	<div>44</div> <div>Ru</div> <div>101.07</div>	<div>45</div> <div>Rh</div> <div>102.91</div>	<div>46</div> <div>Pd</div> <div>106.42</div>	<div>47</div> <div>Ag</div> <div>107.868</div>	<div>48</div> <div>Cd</div> <div>112.411</div>	<div>49</div> <div>In</div> <div>114.818</div>	<div>50</div> <div>Sn</div> <div>118.710</div>	<div>51</div> <div>Sb</div> <div>121.757</div>	<div>52</div> <div>Te</div> <div>127.6</div>	<div>53</div> <div>I</div> <div>126.905</div>	<div>54</div> <div>Xe</div> <div>131.29</div>				
6	<div>55</div> <div>Cs</div> <div>132.905</div>	<div>56</div> <div>Ba</div> <div>137.327</div>	<div>57</div> <div>La</div> <div>138.905</div>	<div>72</div> <div>Hf</div> <div>178.49</div>	<div>73</div> <div>Ta</div> <div>180.948</div>	<div>74</div> <div>W</div> <div>183.84</div>	<div>75</div> <div>Re</div> <div>186.207</div>	<div>76</div> <div>Os</div> <div>190.23</div>	<div>77</div> <div>Ir</div> <div>192.222</div>	<div>78</div> <div>Pt</div> <div>195.084</div>	<div>79</div> <div>Au</div> <div>196.967</div>	<div>80</div> <div>Hg</div> <div>200.59</div>	<div>81</div> <div>Tl</div> <div>204.383</div>	<div>82</div> <div>Pb</div> <div>207.2</div>	<div>83</div> <div>Bi</div> <div>208.980</div>	<div>84</div> <div>Po</div> <div>209</div>	<div>85</div> <div>At</div> <div>210</div>	<div>86</div> <div>Rn</div> <div>222</div>				
7	<div>87</div> <div>Fr</div> <div>223</div>	<div>88</div> <div>Ra</div> <div>226</div>	<div>89</div> <div>Ac</div> <div>227</div>	<div>104</div> <div>Unq</div> <div>261</div>	<div>105</div> <div>Unp</div> <div>262</div>	<div>106</div> <div>Unh</div> <div>263</div>	<div>107</div> <div>Uns</div> <div>264</div>															

**THE SYMBOL.** Shown in the middle of each block directly below the name of the element.  
The color used indicates the physical state of the element under ordinary conditions: black for solids, green for liquids and blue for gases.

**THE ATOMIC WEIGHT.** Shown below the symbol for each element - the atomic weight is shown in black. The values are taken from the official Report on Atomic Weights (J. J. Auer, Chem. Soc. 1913-1914). For elements not listed in the Report the mass number of the longest lived isotopes is shown in brackets.

**THE ATOMIC NUMBER.** Shown in red in the upper left hand corner.

**ELECTRONIC CONFIGURATION.** Shown at the upper right as a group of black numbers. When read downward they indicate the number of electrons normally found in successive energy levels.

58	59	60	61	62	63	64	65	66	67	68	69	70	71
<b>Ce</b>	<b>Pr</b>	<b>Nd</b>	<b>Pm</b>	<b>Sm</b>	<b>Eu</b>	<b>Gd</b>	<b>Tb</b>	<b>Dy</b>	<b>Ho</b>	<b>Er</b>	<b>Tm</b>	<b>Yb</b>	<b>Lu</b>

90	91	92	93	94	95	96	97	98	99	100	101	102	103
<b>Th</b>	<b>Pa</b>	<b>U</b>	<b>Np</b>	<b>Pu</b>	<b>Am</b>	<b>Cm</b>	<b>Bk</b>	<b>Cf</b>	<b>Es</b>	<b>Fm</b>	<b>Md</b>	<b>No</b>	<b>Lr</b>

# Elements Of Old English

**Patrick Vollmar**



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