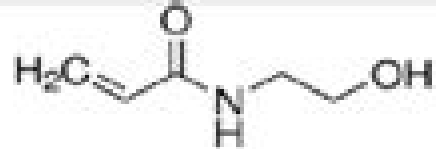
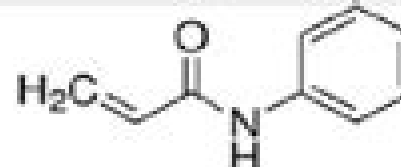


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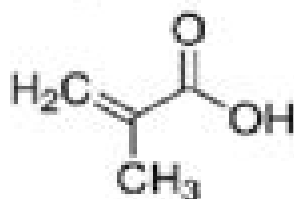
N-hydroxyethyl acrylamide (HEAA)



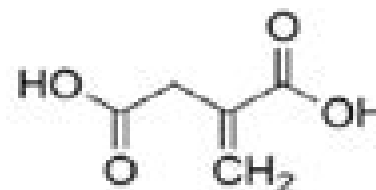
N-phenylacrylamide (PAM)



2-acrylamido-2-methyl-1-propanesulfonic acid (AMPS)



Methacrylic acid (MAA)

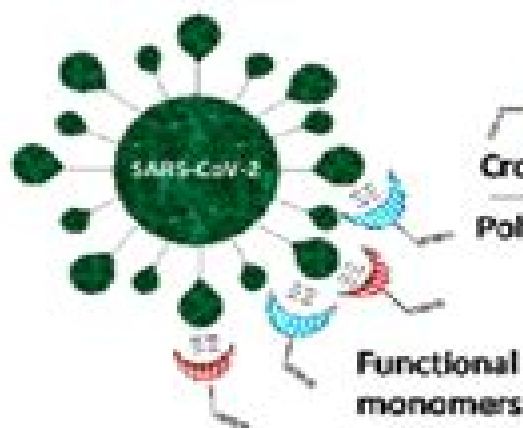


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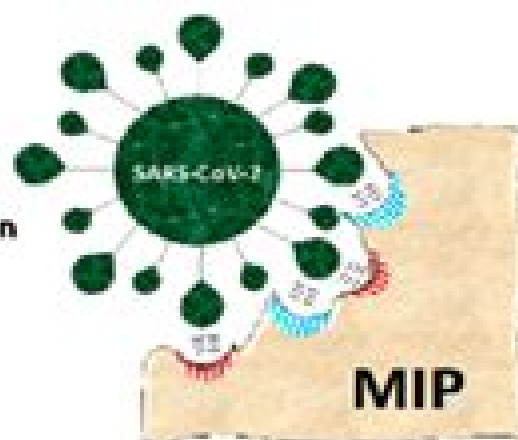


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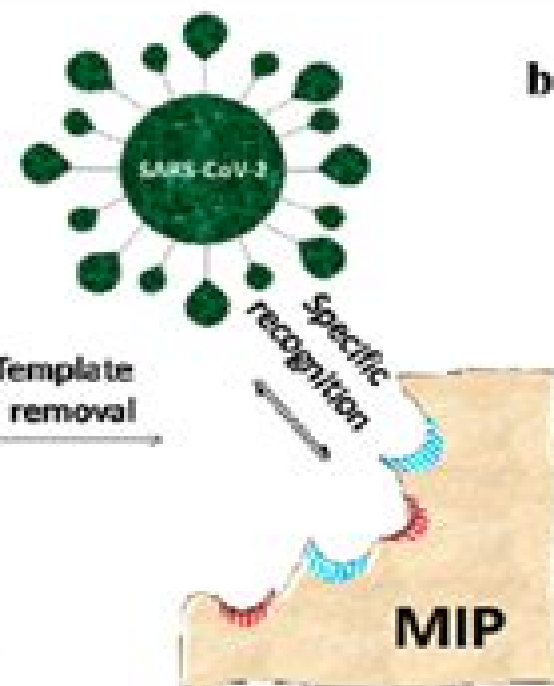
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Polymers have achieved an enviable position as the class of materials having the highest volume of production exceeding that of both metals and ceramics The meteoric rise in the production and utilization of polymers has been due to advances in polymer synthesis which allow the creation of specific and well defined molecular structures to new knowledge concerning the relationships between polymer structure and properties and to an improved understanding of how processing can be used as a tool to develop morphological features which result in desired properties Polymers have truly become engineered materials in every sense of the term Polymer scientists and engineers are forever seeking to modify and improve the properties of synthetic polymeric systems for use in specific applications Towards this end they have often looked to nature for advice on how to design molecules for specific needs An excellent illustration of this is the use of noncovalent bonding ionic hydrogen and van der Waals in lipids proteins and nucleic acids where these noncovalent bonds acting both intra and intermolecularly precisely control the structure and thus the function of the entire system The utilization of ionic bonding in particular in man made polymers has attracted widespread interest in recent years since ionic interactions exert a similar strong influence on the structure and properties of these synthetic systems

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