

Elements of Group IVA or 14 (Elements of Carbon Family, $ns^2 np^2$)

POSITION IN PERIODIC TABLE

Group IVA or 14 of long form of periodic table consists of six elements-carbon (C), silicon (Si), germanium (Ge), tin (Sn), and lead (Pb). These five elements constitute a family known as carbon family. These are p -block elements as the last differentiating electron is accommodated in np shell. These elements have four electrons in their valency shell and thus placed in the IVth group. The elements of this group show similarities as well as gradation in properties with rise of atomic number and thus their inclusion in the same group is justified on the basis of electronic configuration and physical and chemical properties which are described below

1-Electronic configuration : The electronic configuration of the members of this group are:

Element	At.NO.	Electronic configuration	Inert gas core
Carbon - C	6	$1s^2 2s^2 2p^2$	$[\text{He}] 2s^2 2p^2$
Silicon - Si	14	$1s^2 2s^2 2p^6 3s^2 3p^2$	$[\text{Ne}] 3s^2 3p^2$
Germanium - Ge	32	$1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^2$	$[\text{Ar}] 3d^{10} 4s^2 4p^2$
Tin - Sn	50	$1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^{10} 5s^2 5p^2$	$[\text{Kr}] 4d^{10} 5s^2 5p^2$
Lead - Pb	82	$1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^{10} 4f^{14} 5s^2 5p^6 5d^{10} 6s^2 6p^2$	$[\text{Xe}] 4f^{14} 5d^{10} 6s^2 6p^2$

The configurations show that these elements have same number of electrons in the valency shell, i.e., 4 electrons in the valency shell, two of which are in s -orbital while remaining two in p -orbitals. Thus, they have $ns^2 np^2$ configuration, i.e., s -orbital is paired while two p -orbitals are unpaired.



The penultimate shell of carbon contains 2 electrons (saturated), of silicon contains 8 electrons (saturated), of germanium contains 18 electrons (saturated) while Sn and Pb contain 18 electrons (unsaturated) each. This shows why carbon differs from silicon in some respects and these two differ from rest of the members of this group. The close resemblance between the elements is quite striking in the case of elements of the first and second groups, but this becomes less evident in the elements of third group and still less evident in the fourth group.

Elements Compounds Of Group Iva Volume 7 C

Shasha Hu



Elements Compounds Of Group Iva Volume 7 C:

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Table of Contents Elements Compounds Of Group Iva Volume 7 C

1. Understanding the eBook Elements Compounds Of Group Iva Volume 7 C
 - The Rise of Digital Reading Elements Compounds Of Group Iva Volume 7 C
 - Advantages of eBooks Over Traditional Books
2. Identifying Elements Compounds Of Group Iva Volume 7 C
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Elements Compounds Of Group Iva Volume 7 C
 - User-Friendly Interface
4. Exploring eBook Recommendations from Elements Compounds Of Group Iva Volume 7 C
 - Personalized Recommendations
 - Elements Compounds Of Group Iva Volume 7 C User Reviews and Ratings

- Elements Compounds Of Group Iva Volume 7 C and Bestseller Lists
- 5. Accessing Elements Compounds Of Group Iva Volume 7 C Free and Paid eBooks
 - Elements Compounds Of Group Iva Volume 7 C Public Domain eBooks
 - Elements Compounds Of Group Iva Volume 7 C eBook Subscription Services
 - Elements Compounds Of Group Iva Volume 7 C Budget-Friendly Options
- 6. Navigating Elements Compounds Of Group Iva Volume 7 C eBook Formats
 - ePub, PDF, MOBI, and More
 - Elements Compounds Of Group Iva Volume 7 C Compatibility with Devices
 - Elements Compounds Of Group Iva Volume 7 C Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Elements Compounds Of Group Iva Volume 7 C
 - Highlighting and Note-Taking Elements Compounds Of Group Iva Volume 7 C
 - Interactive Elements Elements Compounds Of Group Iva Volume 7 C
- 8. Staying Engaged with Elements Compounds Of Group Iva Volume 7 C
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Elements Compounds Of Group Iva Volume 7 C
- 9. Balancing eBooks and Physical Books Elements Compounds Of Group Iva Volume 7 C
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Elements Compounds Of Group Iva Volume 7 C
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Elements Compounds Of Group Iva Volume 7 C
 - Setting Reading Goals Elements Compounds Of Group Iva Volume 7 C
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Elements Compounds Of Group Iva Volume 7 C
 - Fact-Checking eBook Content of Elements Compounds Of Group Iva Volume 7 C
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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