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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the activity series chart a Al b  $\text{Ba}^{2+}$  c  $\text{Br}_2$  d Ca e  $\text{Ga}^{3+}$  f H<sub>2</sub> g H

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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web an oxidation reduction redox reaction is one in which one or more atoms change oxidation numbers oxidation occurs when an atom's oxidation state becomes more positive indicating that electrons have been lost reduction occurs when an atom's oxidation state becomes less positive indicating that electrons have been gained

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23  $\text{Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$   
oxidation  $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$  reduction  $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$  24 3  $\text{Ag} + \text{Ni}^{3+} \rightarrow \text{Ni} + 3\text{Ag}^+$  oxidation reduction 25  $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$  oxidation reduction 26  $\text{O}_2$

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also 2  $\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   $\text{Sr} \rightarrow \text{Sr}^{2+}$  oxidized reducing agent  $\text{O}_2 \rightarrow \text{O}^{2-}$  reduced ox ag 2  $\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$   $\text{Li} \rightarrow \text{Li}^+$  oxidized red ag  $\text{S} \rightarrow \text{S}^{2-}$  reduced ox ag 3  $2\text{Cs} + \text{Br}_2$

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