

CHM 102

LECTURER: Dr. (Mrs). Tovide O.O

Elements of Group IVA or 14 (Elements of Carbon Family, $ns^2 np^2$)

POSITION IN PERIODIC TABLE

Group IVA or 14 of long form of periodic table consists of six elements-carbon (C), silicon (Si), germanium (Ge), tin (Sn), and lead (Pb). These five elements constitute a family known as carbon family. These are p -block elements as the last differentiating electron is accommodated in np shell. These elements have four electrons in their valency shell and thus placed in the IVth group. The elements of this group show similarities as well as gradation in properties with rise of atomic number and thus their inclusion in the same group is justified on the basis of electronic configuration and physical and chemical properties which are described below

1-Electronic configuration : The electronic configuration of the members of this group are:

Element	At.NO.	Electronic configuration	Inert gas core
Carbon - C	6	$1s^2 2s^2 2p^2$	$[\text{He}] 2s^2 2p^2$
Silicon - Si	14	$1s^2 2s^2 2p^6 3s^2 3p^2$	$[\text{Ne}] 3s^2 3p^2$
Germanium - Ge	32	$1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^2$	$[\text{Ar}] 3d^{10} 4s^2 4p^2$
Tin - Sn	50	$1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^{10} 5s^2 5p^2$	$[\text{Kr}] 4d^{10} 5s^2 5p^2$
Lead - Pb	82	$1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^{10} 4f^{14} 5s^2 5p^6 5d^{10} 6s^2 6p^2$	$[\text{Xe}] 4f^{14} 5d^{10} 6s^2 6p^2$

The configurations show that these elements have same number of electrons in the valency shell, i.e., 4 electrons in the valency shell, two of which are in s -orbital while remaining two in p -orbitals. Thus, they have $ns^2 np^2$ configuration, i.e., s -orbital is paired while two p -orbitals are unpaired.



The penultimate shell of carbon contains 2 electrons (saturated), of silicon contains 8 electrons (saturated), of germanium contains 18 electrons (saturated) while Sn and Pb contain 18 electrons (unsaturated) each. This shows why carbon differs from silicon in some respects and these two differ from rest of the members of this group. The close resemblance between the elements is quite striking in the case of elements of the first and second groups, but this becomes less evident in the elements of third group and still less evident in the fourth group.

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