



SCIENCE

1. The first part of the experiment was to determine the effect of temperature on the rate of reaction. The temperature was varied from 10°C to 30°C in increments of 5°C. The rate of reaction was measured by the time taken for a fixed volume of gas to be evolved.

2. The second part of the experiment was to determine the effect of concentration on the rate of reaction. The concentration of the reactants was varied from 0.1 M to 0.5 M in increments of 0.1 M. The rate of reaction was measured by the time taken for a fixed volume of gas to be evolved.

3. The third part of the experiment was to determine the effect of surface area on the rate of reaction. The surface area of the reactants was varied by using different sizes of particles. The rate of reaction was measured by the time taken for a fixed volume of gas to be evolved.

4. The fourth part of the experiment was to determine the effect of a catalyst on the rate of reaction. The effect of a catalyst was tested by measuring the rate of reaction with and without the catalyst. The rate of reaction was measured by the time taken for a fixed volume of gas to be evolved.

5. The fifth part of the experiment was to determine the effect of pressure on the rate of reaction. The pressure was varied from 1 atm to 2 atm. The rate of reaction was measured by the time taken for a fixed volume of gas to be evolved.

6. The sixth part of the experiment was to determine the effect of the nature of the reactants on the rate of reaction. The rate of reaction was measured by the time taken for a fixed volume of gas to be evolved.

7. The seventh part of the experiment was to determine the effect of the nature of the catalyst on the rate of reaction. The rate of reaction was measured by the time taken for a fixed volume of gas to be evolved.

8. The eighth part of the experiment was to determine the effect of the concentration of the catalyst on the rate of reaction. The rate of reaction was measured by the time taken for a fixed volume of gas to be evolved.

9. The ninth part of the experiment was to determine the effect of the concentration of the reactants on the rate of reaction. The rate of reaction was measured by the time taken for a fixed volume of gas to be evolved.

10. The tenth part of the experiment was to determine the effect of the concentration of the catalyst on the rate of reaction. The rate of reaction was measured by the time taken for a fixed volume of gas to be evolved.

11. The eleventh part of the experiment was to determine the effect of the concentration of the reactants on the rate of reaction. The rate of reaction was measured by the time taken for a fixed volume of gas to be evolved.



12. The twelfth part of the experiment was to determine the effect of the concentration of the reactants on the rate of reaction. The rate of reaction was measured by the time taken for a fixed volume of gas to be evolved.



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