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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of hcl because this is a strong acid the ionization is complete and the hydronium ion molarity

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 $we b \ titration \ experiment \ in \ the \ neutralization \ of \ hydrochloric \ acid \ by \ sodium \ hydroxide \ the \ mole \ ratio \ of \ acid \ to \ base \ is \ 1 \ 1$

hcl aq naoh aq nacl aq h2o l hcl a q

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web oct 15 2022 it should be noted that the ph scale does extend beyond 0 and 14 strong laboratory acids typically have ph values less than 0 negative ph values and strong

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web one type of titration uses a neutralization reaction in which an acid and a base react to produce a salt and water in equation 1 the acid is hcl hydrochloric acid and the base