

CHAPTER NINE

ELECTROCHEMISTRY

Electrochemistry is the branch of chemistry that deals with chemical changes that are produced by passing electric current into a chemical system or the generation of electricity by spontaneous chemical reactions.

MODES OF CONDUCTION OF SUBSTANCES

The substances that allow electric current to flow through them are called conductors. Those that do not allow electric current to flow through them are called non-conductors or insulators. The conductors are classified into two type i.e. *metallic conductors* and *electrolytic conductors*.

1. Metallic conductors

Metals or alloys allow electric current to flow through them without undergoing any chemical change. The metals conduct electric current by use of *free mobile electrons (delocalised electrons)*. These include copper, silver, aluminium and others. This mode of conduction is called *electronic conduction*.

2. Electrolytic conductors

These include *aqueous solutions of electrolytes* or molten electrolytes. They are compounds and chemical changes occur when they conduct electricity. The electrolytes conduct electric current by use of *free mobile ions*. Examples include dilute hydrochloric acid, aqueous sodium chloride, molten lead(II) bromide and others. This mode of conduction is called *electrolytic conduction*.

ELECTROLYSIS

Electrolysis is the decomposition of a compound in its molten state or in solution by passage of an electric current through it.

or

The overall chemical change that occurs at the electrodes when an electric current passes through a fused electrolyte or its aqueous solution

The electrolyte, solvent or electrode may all be involved in the electrolytic reaction

The common terms used in electrolysis

1. Electrolytes

These are compounds which conduct electric current in solution or in molten state and are decomposed by it.

A non-electrolyte is a substance which does not conduct electric current in molten state or aqueous solution and cannot be decomposed by electric current. For example sucrose

Electrolytes can be salts, acids or bases.

Electrolytes can be classified as *strong* or *weak*.

Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34

M Lipman



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Table of Contents Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34

1. Understanding the eBook **Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34**
 - The Rise of Digital Reading **Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34**
 - Advantages of eBooks Over Traditional Books
2. Identifying **Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34**

- User-Friendly Interface
- 4. Exploring eBook Recommendations from Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34
 - Personalized Recommendations
 - Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34 User Reviews and Ratings
 - Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34 and Bestseller Lists
- 5. Accessing Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34 Free and Paid eBooks
 - Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34 Public Domain eBooks
 - Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34 eBook Subscription Services
 - Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34 Budget-Friendly Options
- 6. Navigating Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34 eBook Formats
 - ePub, PDF, MOBI, and More
 - Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34 Compatibility with Devices
 - Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34
 - Highlighting and Note-Taking Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34
 - Interactive Elements Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34

8. Staying Engaged with Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34
9. Balancing eBooks and Physical Books Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34
 - Setting Reading Goals Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34
 - Fact-Checking eBook Content of Electrochemistry The Interfacing Science Studies In Physical And Theoretical Chemistry Volume 34
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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