

Victor Hugo's Works.

30.

THE LIFE, TIMES, AND
WORKS OF VICTOR HUGO.

VOLUME II.

—
Large Paper.

Guide To The Life Times And Works Of Victor Hugo

Christian G. Meyer



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web in today s experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product expression describes in mathematical terms the equilibrium

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the

solution $\text{Ca}(\text{OH})_2$ à Ca^{2+} $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ $2 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide K_{sp} using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the K_{sp} of a slightly soluble basic salt and how solubility changes with temperature K_{sp} is another form of equilibrium expression that deals with solubility the reaction that will take place is $\text{Ca}(\text{OH})_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web the chemical equation used was $\text{Ca}(\text{OH})_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product was determined to be K_{sp} Ca^{2+} OH⁻ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $\text{Ca}_3(\text{PO}_4)_2 \rightleftharpoons 3\text{Ca}^{2+} + 2\text{PO}_4^{3-}$ K_{sp} Ca^{2+} PO_4^{3-}

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web lab report 8 determining K_{sp} of calcium hydroxide the purpose of this was to titrate a saturated $\text{Ca}(\text{OH})_2$ solution with a standard HCl solution and find the K_{sp} or the OH⁻ concentration was found and then the concentration was found and used for this experiment a titration of calcium hydroxide with hydrochloric acid

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web calculate the ksp for calcium hydroxide $\text{ksp Ca(OH)}_2 = 5 \times 10^{-6}$ find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error accepted value for ksp for sodium hydroxide is 5×10^{-6}

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