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was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
 one mole of HCl would be fully neutralized by one mole of NaOH

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown

acid or base is strong or weak pka of an unknown acid or pkb of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4$ aq 2NaOH aq $\text{Na}_2\text{C}_2\text{O}_4$ $4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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