

Table 8.1: Fundamental physical formats, constants, and relationships between the units of the elementary quantities count, amount, and charge.

Format	Name	Definitions	Numerical value	Unit *
\underline{N}	count, molecular #	count $N_X = N \cdot U_X$	\triangle MU = x	
\underline{n}	amount, molar	amount $n_X = N_X \cdot N_A^{-1}$	\square MU = mol	
\underline{q}	charge, electrical	charge $Q_{eX} = n_X \cdot z_X \cdot F$	\circledast MU = C	
\underline{N}	Boltzmann constant §	$k = f \cdot e = R / N_A = 1.380\,649 \cdot 10^{-23}$	$\text{J} \cdot \text{x}^{-1} \cdot \text{K}^{-1}$	
\underline{n}	gas constant	$R = f \cdot F = k \cdot N_A = 8.314\,462\,618$	$\text{J} \cdot \text{mol}^{-1} \cdot \text{K}^{-1}$	
\underline{q}	electromotive constant §	$f = k / e = R / F = 8.617\,333\,262 \cdot 10^{-5}$	$\text{J} \cdot \text{C}^{-1} \cdot \text{K}^{-1}$	
$\underline{N}/\underline{n}$	Avogadro constant §	$N_A = R / k = F / e = 6.022\,140\,76 \cdot 10^{23}$	$\text{x} \cdot \text{mol}^{-1}$	
$\underline{q}/\underline{N}$	elementary charge §,®	$e = k / f = F / N_A = 1.602\,176\,634 \cdot 10^{-19}$	$\text{C} \cdot \text{x}^{-1}$	
$\underline{q}/\underline{n}$	Faraday constant	$F = R / f = e \cdot N_A = 96\,485.332\,12$	$\text{C} \cdot \text{mol}^{-1}$	

* The motive quantity with motive unit MU defines the physical format.

® Count N_X is the number N of elementary entities U_X with the abstract elementary unit [x] (Gnaiger 2020).

§ Redefinition of SI base units came into force on 2019-05-20; Bureau International des Poids et Mesures (2019) The International System of Units (SI), 9th edition.

§ A name or symbol was not found in the literature for the electromotive constant f introduced here.

® Elementary charge $e \triangleq Q_{dp^+} / N_{p^+} = Q_{dp^+}$ is charge per proton count or charge per elementary proton U_{p^+} .

Electrochemical Constants

Rüdiger Memming

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Electrochemical Constants:

Electrochemical Constants United States. National Bureau of Standards, 1953 **Handbook of Electrochemistry** Cynthia G. Zoski, 2007-02-07 Electrochemistry plays a key role in a broad range of research and applied areas including the exploration of new inorganic and organic compounds biochemical and biological systems corrosion energy applications involving fuel cells and solar cells and nanoscale investigations The Handbook of Electrochemistry serves as a source of electrochemical information providing details of experimental considerations representative calculations and illustrations of the possibilities available in electrochemical experimentation The book is divided into five parts Fundamentals Laboratory Practical Techniques Applications and Data The first section covers the fundamentals of electrochemistry which are essential for everyone working in the field presenting an overview of electrochemical conventions terminology fundamental equations and electrochemical cells experiments literature textbooks and specialized books Part 2 focuses on the different laboratory aspects of electrochemistry which is followed by a review of the various electrochemical techniques ranging from classical experiments to scanning electrochemical microscopy electrogenerated chemiluminescence and spectroelectrochemistry Applications of electrochemistry include electrode kinetic determinations unique aspects of metal deposition and electrochemistry in small places and at novel interfaces and these are detailed in Part 4 The remaining three chapters provide useful electrochemical data and information involving electrode potentials diffusion coefficients and methods used in measuring liquid junction potentials serves as a source of electrochemical information includes useful electrochemical data and information involving electrode potentials diffusion coefficients and methods used in measuring liquid junction potentials reviews electrochemical techniques incl scanning electrochemical microscopy electrogenerated chemiluminescence and spectroelectrochemistry **Electrochemical Constants** United States. National Bureau of Standards, 1953

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Electrochemical Methods of Detoxification for Medical Use V. I. Sergienko, Yu B. Vasiliev, 1989 Electrochemical Constants ,1951

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