

BOLTZMANN ENTROPY: PROBABILITY AND INFORMATION

We have presented first an axiomatic derivation of Boltzmann entropy on the basis of two axioms consistent with two basic properties of thermodynamic entropy. We have then studied the relationship between Boltzmann entropy and information along with its physical significance. PACS : 89.70, + C.

Keywords: Boltzmann entropy, thermodynamic probability, axiomatic derivation, information, statistical equilibrium.

1. INTRODUCTION

The concept of entropy was first introduced in thermodynamics by Clausius through the second law of thermodynamics. It was Boltzmann who first provided the statistical analogue of thermodynamic entropy linking the concept of entropy with molecular disorder or chaos [1]. The concept of entropy was later developed with the advent of statistical mechanics and information theory [3]. Boltzmann entropy is a building block to the foundation of statistical mechanics and the literature on Boltzmann entropy is vast. Inspite of this, there is scope for further extension of the existing methods and principles. In the present paper we have provided first an axiomatic derivation of Boltzmann entropy on the basis of two axioms consistent with the additivity and the increasing law of thermodynamic entropy. We have then introduced the probability distribution of a macrostate through the concept of prior probabilities. This has led to an interesting relation connecting negentropy and information. The physical significance of this relation in the characterization of statistical equilibrium has been investigated.

2. THERMODYNAMIC PROBABILITY AND BOLTZMANN ENTROPY

Boltzmann entropy is defined by [1]

$$S = k \ln W \quad \dots (2.1)$$

where k is the thermodynamic unit of the measurement of the entropy and is the Boltzmann constant, W called the thermodynamic probability or statistical weight is the total number of microscopic states or complexions compatible with the macroscopic state of the system. Following Carnot [4], we shall, however, call W as the degree of disorder or simply disorder of the system. We can thus restate Boltzmann principle as

$$\text{Entropy} = k \ln(\text{Disorder}) \quad \dots (2.2)$$

The entropy S being a monotonic increasing function of disorder, can itself be considered as a measure of disorder of the system. Boltzmann entropy thus provides a measure of the disorder of the distribution of the states over permissible microstates [4]. Let us derive the expression (2.1) of entropy axiomatically on the

basis of two fundamental properties of thermodynamic entropy. We make the following axioms :

Axiom (1) :

The entropy $S(W)$ of system is a positive increasing function of the disorder W , that is,

$$S(W) \leq S(W+1) \quad \dots (2.3)$$

This is motivated by the increasing law of entropy in thermodynamics.

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