

Electrostatics of Atoms & Molecules

We have considered how *Coulombic* attractions and repulsions help to organize **electrons** in atoms and ions.

We now look at *Coulombic* attractions between **atoms** and **molecules**...

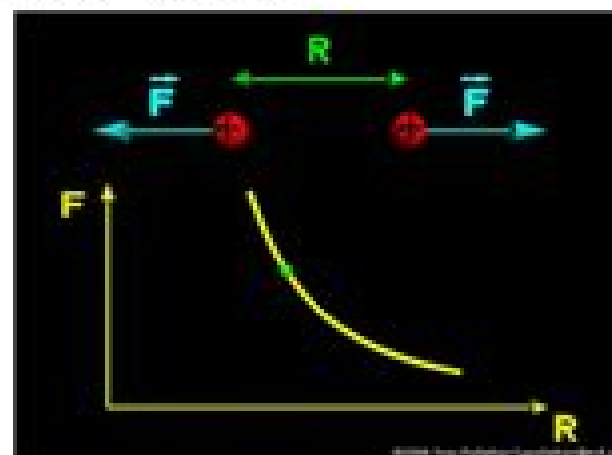
Coulomb's Law has survived every experimental test...

...no exceptions have ever been found.

It holds deep within the atom, correctly describing the force between the (+)vely charged nucleus and each (-)vely charged extra-nuclear electron.

It holds up where classical Newtonian mechanics breaks down!

Coulomb's Law:
$$F = C \frac{q_1 q_2}{r^2}$$



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Emilie Sanchez



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