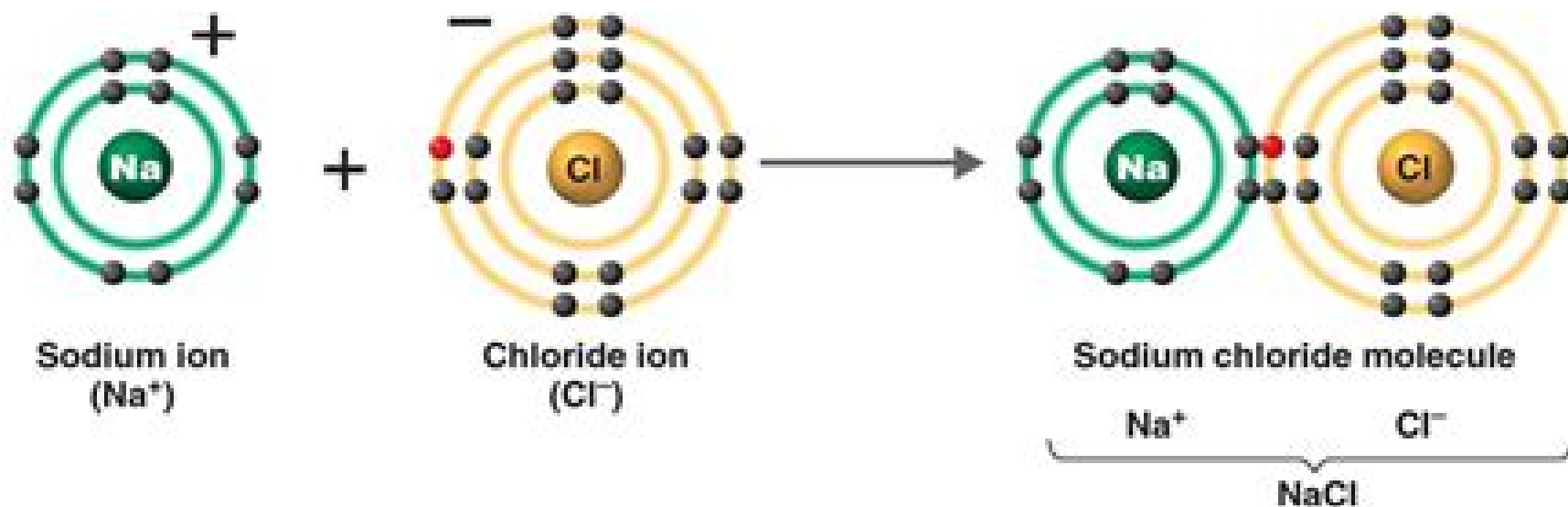


(a) A sodium atom (Na) loses one electron to an electron acceptor and forms a sodium ion (Na⁺). A chlorine atom (Cl) accepts one electron from an electron donor to become a chloride ion (Cl⁻).



(b) The sodium and chloride ions are attracted because of their opposite charges and are held together by an ionic bond to form a molecule of sodium chloride.

Electronic Structure Chemical Bonding

Donald Keith Sebera



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Electronic Structure and the Properties of Solids Walter A. Harrison, 2012-03-08 This text offers basic understanding of the electronic structure of covalent and ionic solids simple metals transition metals and their compounds also explains how to calculate dielectric conducting bonding properties

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Electronic Structure and Properties of Transition Metal Compounds Isaac B. Bersuker, 2010-12-01 With more than 40% new and revised materials this second edition offers researchers and students in the field a comprehensive understanding of fundamental molecular properties amidst cutting edge applications Including 70 Example Boxes and summary notes questions exercises problem sets and illustrations in each chapter this publication is also suitable for use as a textbook for advanced undergraduate and graduate students Novel material is introduced in description of multi orbital chemical bonding spectroscopic and magnetic properties methods of electronic structure calculation and quantum classical modeling for organometallic and metallobiochemical systems This is an excellent reference for chemists researchers and teachers and advanced undergraduate and graduate students in inorganic coordination and organometallic chemistry

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a solid with specified properties This book is intended to acquaint the reader with established principles of crystallography and cohesive forces that are needed to address the fundamental relationship between the composition structure and bonding Starting with an introduction to periodic trends the book discusses crystal structures and the various primary and secondary bonding types and finishes by describing a number of models for predicting phase stability and structure Containing a large number of worked examples exercises and detailed descriptions of numerous crystal structures this book is primarily intended as an advanced undergraduate or graduate level textbook for students of materials science It will also be useful to scientists and engineers who work with solid materials

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Electronic Structure N. D. Epiotis, 2012-12-06 Theoretical Models of Chemical Bonding Zvonimir B. Maksic, 2012-12-06

The renowned theoretical physicist Victor F Weisskopf rightly pointed out that a real understanding of natural phenomena implies a clear distinction between the essential and the peripheral Only when we reach such an understanding that is to say when we are able to separate the relevant from the irrelevant will the phenomena no longer appear complex but intellectually transparent This statement which is generally valid reflects the very essence of modelling in the quantum theory of matter on the molecular level in particular Indeed without theoretical models one would be swamped by too many details embodied in intricate accurate molecular wavefunctions Further physically justified simplifications enable studies of the otherwise intractable systems and or phenomena Finally a lack of appropriate models would leave myriads of raw experimental data totally unrelated and incomprehensible The present series of books dwells on the most important models of chemical bonding and on the variety of its manifestations In this volume the electronic structure and properties of molecules are considered in depth Particular attention is focused on the nature of intramolecular interactions which in turn are revealed by various types of molecular spectroscopy Emphasis is put on the conceptual and interpretive aspects of the theory in line with the general philosophy adopted in the series **Research Grants Index** National Institutes of Health (U.S.). Division of Research

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