Hence, FIRST LAW of THERMODYNAMICS (applied to a closed system)

- The change in internal energy of a closed system is the energy that enters or leaves the system through boundaries as heat or work, i.e.
- ∆U = q +w
 - $> \Delta U = U_1 U_1$
 - q heat applied to system
 - W work done on system
 - When energy leaves the system, ∆U = -ve i.e. decrease internal energy
- When energy enter the system, ∆U = +ve i.e. added to internal energy

Elements Of Applied Thermodynamics

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