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 $HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$
 one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 \text{ aq } 2\text{NaOH aq Na}_2\text{C}_2\text{O}_4 \cdot 4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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