

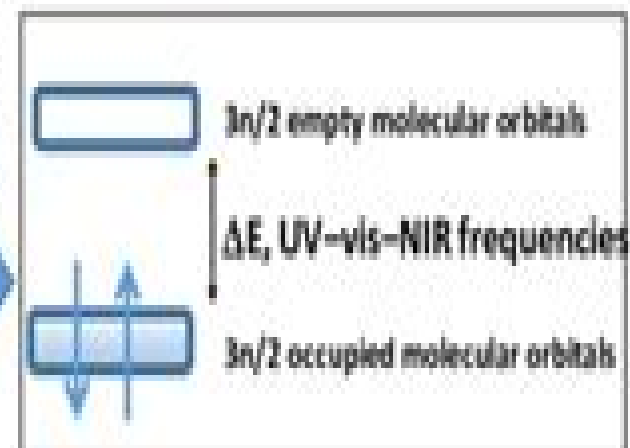


Molecules with alternating singles and double bonds, X = Hydrogen; Heteroatom; conjugated group, cyclic aromatic group. Y = Carbon; heteroatom, cyclic aromatic group.

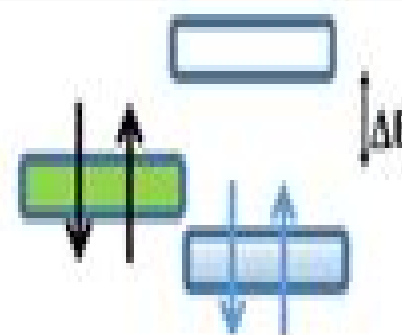
p orbitals along the backbone, each one contains one electron

Overlapping of  $3n$  atomic p orbitals generates  $3n$  molecular orbitals,  $3n/2$  are occupied by  $3n$  electrons,  $3n/2$  are empty

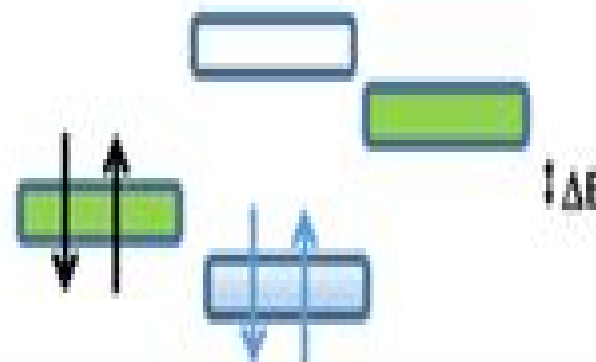
X, Y = Hydrogen, Carbon



X or Y = heteroatom: no bonding orbitals (green colored), containing electron couples, sit between the  $n$  occupied and  $n$  empty orbitals, lowering the  $\Delta E$ .



X = conjugated group: orbitals containing electron couples and empty orbitals (green color) sit between the  $n$  occupied and  $n$  empty orbitals, lowering the  $\Delta E$ .



# Electronic Properties Of Conjugated Polymers

**Anthony S. Fauci**



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