



SCIENCE

1. The first part of the experiment was to determine the effect of temperature on the rate of reaction. The temperature was varied from 10°C to 30°C in increments of 5°C. The rate of reaction was measured by the time taken for a fixed volume of gas to be evolved.

2. The second part of the experiment was to determine the effect of concentration on the rate of reaction. The concentration of the reactants was varied from 0.1 M to 0.5 M in increments of 0.1 M.

3. The third part of the experiment was to determine the effect of surface area on the rate of reaction. The surface area of the reactants was varied by using different sizes of particles.

4. The fourth part of the experiment was to determine the effect of a catalyst on the rate of reaction. The catalyst was varied from none to a fixed amount of a specific catalyst.

5. The fifth part of the experiment was to determine the effect of pressure on the rate of reaction. The pressure was varied from 1 atm to 5 atm in increments of 1 atm.

6. The sixth part of the experiment was to determine the effect of solvent on the rate of reaction. The solvent was varied from water to a fixed amount of a specific solvent.

7. The seventh part of the experiment was to determine the effect of pH on the rate of reaction. The pH was varied from 1 to 14 in increments of 1.

8. The eighth part of the experiment was to determine the effect of ionic strength on the rate of reaction. The ionic strength was varied from 0.1 M to 0.5 M in increments of 0.1 M.

9. The ninth part of the experiment was to determine the effect of dielectric constant on the rate of reaction. The dielectric constant was varied from 1 to 10 in increments of 1.

10. The tenth part of the experiment was to determine the effect of viscosity on the rate of reaction. The viscosity was varied from 0.1 cP to 1.0 cP in increments of 0.1 cP.

11. The eleventh part of the experiment was to determine the effect of surface tension on the rate of reaction. The surface tension was varied from 0.1 dyne/cm to 1.0 dyne/cm in increments of 0.1 dyne/cm.

12. The twelfth part of the experiment was to determine the effect of adsorption on the rate of reaction. The adsorption was varied from 0.1 g/g to 1.0 g/g in increments of 0.1 g/g.

13. The thirteenth part of the experiment was to determine the effect of desorption on the rate of reaction. The desorption was varied from 0.1 g/g to 1.0 g/g in increments of 0.1 g/g.

14. The fourteenth part of the experiment was to determine the effect of diffusion on the rate of reaction. The diffusion coefficient was varied from 0.1 cm²/s to 1.0 cm²/s in increments of 0.1 cm²/s.

15. The fifteenth part of the experiment was to determine the effect of convection on the rate of reaction. The convection coefficient was varied from 0.1 cm/s to 1.0 cm/s in increments of 0.1 cm/s.



16. The sixteenth part of the experiment was to determine the effect of radiation on the rate of reaction. The radiation dose was varied from 0.1 Mrad to 1.0 Mrad in increments of 0.1 Mrad.

17. The seventeenth part of the experiment was to determine the effect of magnetic field on the rate of reaction. The magnetic field strength was varied from 0.1 T to 1.0 T in increments of 0.1 T.

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- Chapter 2: Essential Elements of Hbj Science Level 4
- Chapter 3: Hbj Science Level 4 in Everyday Life
- Chapter 4: Hbj Science Level 4 in Specific Contexts
- Chapter 5: Conclusion

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